

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS General Certificate of Education Ordinary Level

COMBINED SCIENCE 5129/01

Paper 1 Multiple Choice May/June 2008

1 hour

Additional Materials: Multiple Choice Answer Sheet

Soft clean eraser

Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A**, **B**, **C** and **D**.

Choose the one you consider correct and record your choice in soft pencil on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

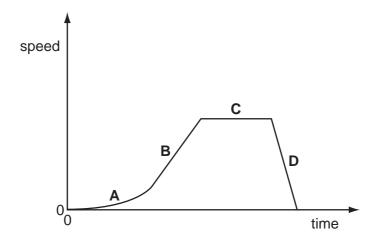
A copy of the Periodic Table is printed on page 20.



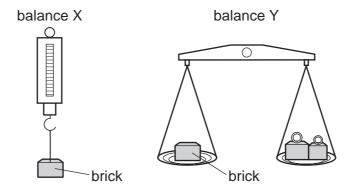
International Examinations

1 The diagram shows a speed-time graph for an object.

Which section of the graph shows this object moving with constant speed?



2 A brick is placed on a newton balance X and then on a beam balance Y.

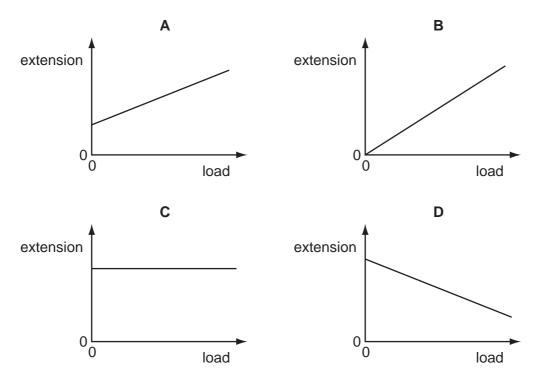


What is measured by each balance?

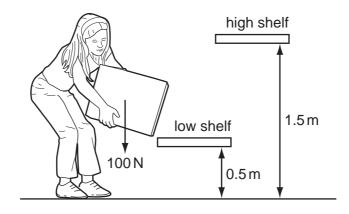
	balance X	balance Y		
Α	mass	mass		
B mass		weight		
С	weight	mass		
D	weight	weight		

3 A student adds different loads to the end of a spring. She finds the extension in each case and plots a graph of extension against load.

Which is the correct graph?



4 The diagram shows a girl lifting a box of weight 100 N from a low shelf to a high shelf.



How much work is done by the girl?

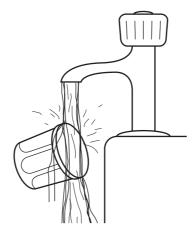
A 50 J

B 100 J

C 150 J

D 200 J

5 A person cannot unscrew the lid of a pot of jam. He finds that the metal lid can be unscrewed after it has been held under hot, running water for a few seconds.



Why is this?

- **A** The air pressure in the jar falls.
- **B** The glass expands.
- C The jam melts.
- **D** The metal lid expands.
- 6 A wave has a frequency of 30 000 Hz and a speed of 1500 m/s.

What is the wavelength?

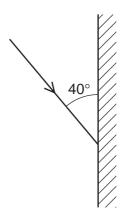
A 0.05 m **B** 0

B 0.50 m

C 20 m

D 200 m

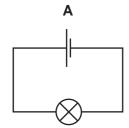
7 The diagram shows a single ray of light being directed at a plane mirror.

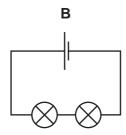


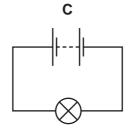
What are the angles of incidence and reflection?

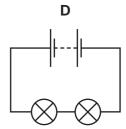
	angle of incidence	angle of reflection
Α	40°	40°
В	40°	50°
С	50°	40°
D	50°	50°

- 8 An electric current in a metal wire involves the movement of
 - A atoms.
 - B electrons.
 - C molecules.
 - **D** protons.
- **9** Which circuit contains the brightest lamp?







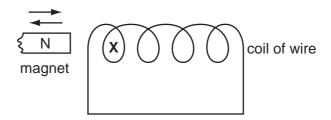


10 A potential difference of 4 V drives a current of 3 A through a resistor.

How much electrical energy is converted into heat during 10s?

- **A** 12 J
- **B** 30 J
- **C** 40 J
- **)** 120 J

11 The diagram shows the north pole of a magnet moved into, and out of, a coil of wire.



What describes the poles produced in the coil at **X** by the movement of the magnet?

	north pole in	north pole out
Α	N	N
В	N	S
С	S	N
D	S	S

12 A nuclide of the element plutonium is $^{242}_{94} \mathrm{Pu}$.

What is the number of neutrons in its nucleus?

- **A** 336
- **B** 242
- **C** 148
- **D** 94

13 The radioactive decay of a nuclide is represented by the equation below.

$$^{234}_{90}$$
Th $ightarrow$ $^{234}_{91}$ Pa + emitted particle

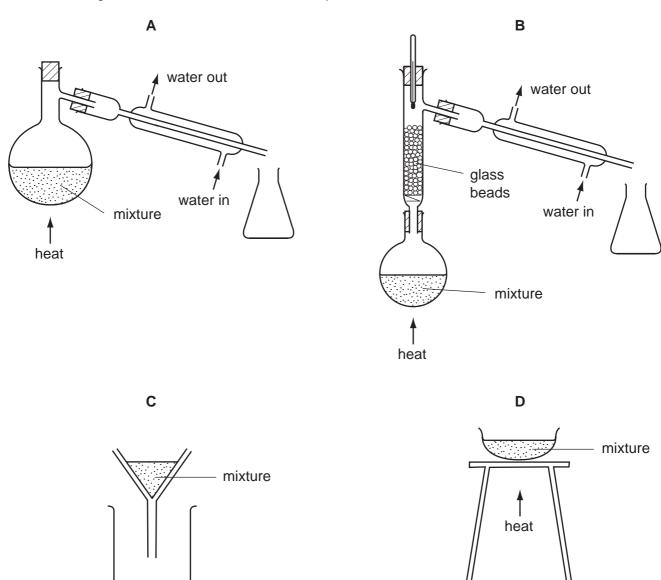
Which type of particle is emitted during the decay shown?

- A alpha-particle
- B beta-particle
- **C** neutron
- **D** proton

14 Substance X melts at 53 °C and boils at 100 °C.

It does not dissolve in water.

Which diagram shows the method used to separate \boldsymbol{X} from a mixture of \boldsymbol{X} and water?



15 If two neutral atoms are isotopes of the same element, they both have the same number of

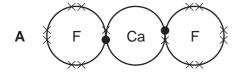
- 1 particles in the nucleus.
- 2 electrons.
- 3 neutrons.
- 4 protons.

Which statements are correct?

- **A** 1, 2 and 3
- **B** 1 and 3 only
- **C** 2 and 4
- **D** 4 only

16 Which diagram shows the electron arrangement in calcium fluoride?

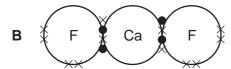
Only the outermost electrons of each ion are shown.

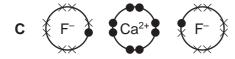


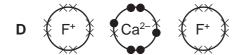
key

• = electrons from calcium

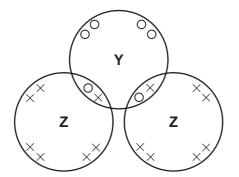
 \times = electrons from fluorine







17 The diagram shows the arrangement of electrons in a molecule of compound YZ₂.



key

O = outer electron of a Y atom

 \times = outer electron of a **Z** atom

What are elements Y and Z?

	Y	Z
Α	calcium chlorine	
В	carbon	oxygen
С	oxygen	hydrogen
D	sulphur	chlorine

18 25.0 g of hydrated copper(II) sulphate crystals are heated to produce anhydrous copper(II) sulphate and water vapour.

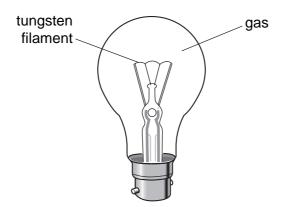
$$CuSO_4$$
 . $5H_2O \rightarrow CuSO_4 + 5H_2O$

What mass of anhydrous copper(II) sulphate is formed? [CuSO₄ = 160; H_2O = 18.]

- **A** 9.0 g
- **B** 16.0 g
- **C** 22.5 g
- **D** 25.0 g

- 19 Which compound is an amphoteric oxide?
 - A calcium oxide
 - B copper(II) oxide
 - C sulphur dioxide
 - **D** zinc oxide

20 Which gas is present in the light bulb?



- **A** argon
- B carbon dioxide
- C nitrogen
- **D** oxygen

21 Water is formed when hydrogen is passed over the heated oxide of metal X.

No water is formed when hydrogen is passed over the heated oxide of metal Y.

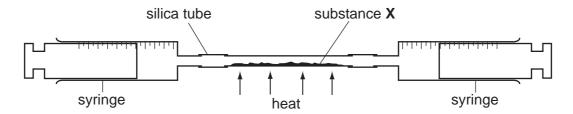
What is the order of reactivity of hydrogen, metal **X** and metal **Y**?

	most reactive		least reactive
Α	hydrogen	X	Y
В	x	hydrogen	Y
С	x	Y	hydrogen
D	Y	hydrogen	X

22 Which metal is used for galvanising?

- A aluminium
- **B** copper
- C iron
- **D** zinc

23 The apparatus shown is used to measure the percentage by volume of oxygen in the air.



What is substance X?

- A anhydrous copper(II) sulphate
- B calcium oxide
- C carbon
- **D** copper
- **24** Ammonium sulphate, $(NH_4)_2SO_4$, is sometimes added to soil to provide an element that is important for plant growth.

What is this element?

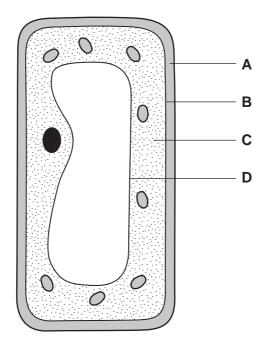
- A hydrogen
- **B** nitrogen
- C oxygen
- **D** sulphur
- 25 In which of the following are all the compounds members of the same homologous series?
 - A CH_4 C_2H_6 C_3H_6
 - $\mathbf{B} \quad \mathsf{CH}_4 \qquad \mathsf{C}_2\mathsf{H}_6 \qquad \mathsf{C}_3\mathsf{H}_8$
 - ${f C} \quad {f C}_2{f H}_4 \quad {f C}_3{f H}_6 \quad {f C}_4{f H}_{10}$
 - **D** C_3H_4 C_3H_6 C_3H_8
- **26** Four of the products of the fractional distillation of petroleum are diesel oil, gasoline, kerosene and lubricating oil.

In which order do they distil off, lowest boiling point first?

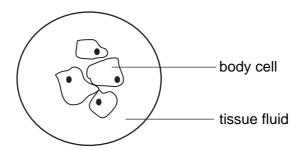
- **A** diesel oil \rightarrow gasoline \rightarrow kerosene \rightarrow lubricating oil
- **B** gasoline \rightarrow kerosene \rightarrow diesel oil \rightarrow lubricating oil
- **C** gasoline \rightarrow kerosene \rightarrow lubricating oil \rightarrow diesel oil
- **D** kerosene \rightarrow gasoline \rightarrow diesel oil \rightarrow lubricating oil

- 27 What does **not** happen in the combustion of pure ethane in a plentiful supply of air?
 - A a smoky flame is seen
 - **B** carbon dioxide is produced
 - C energy is released
 - **D** water is produced
- 28 The diagram shows a plant cell.

Which structure is the cell membrane?



29 The diagram shows a group of body cells surrounded by tissue fluid.



Which conditions cause the body cells to take in water?

	concentration of water in the tissue fluid	concentration of water in the cytoplasm of body cells
Α	high	high
В	high	low
С	low	high
D	low	low

- **30** Four types of cell found in the leaf of a green plant are listed below.
 - 1 epidermal cells (not including guard cells)
 - 2 guard cells
 - 3 palisade mesophyll cells
 - 4 spongy mesophyll cells

Which cells contain chloroplasts?

- A 1 and 2 only
- B 2 and 3 only
- **C** 2, 3 and 4 only
- **D** 1, 2, 3 and 4

31 The diagram represents stages in the breakdown of starch to maltose by the enzyme amylase.



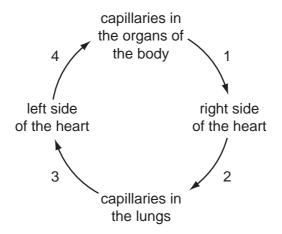
What are the correct labels?

	amylase	maltose	starch
Α	Р	S	Q
В	Q	R	S
С	R	Q	Р
D	S	Р	R

32 A young plant is dug up and then re-planted. Later, the plant wilts.

What causes this?

- A The leaves lose less water.
- **B** The roots cannot take up mineral ions.
- C The stomata close.
- **D** The surface area of the roots is reduced.
- **33** The diagram shows the direction of blood flow in the human body.



At which stages does the blood contain the most oxygen?

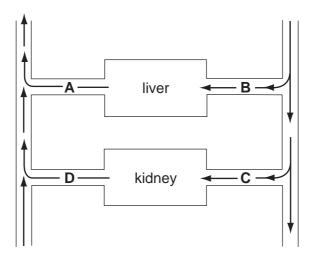
A 1 and 2 **B** 2 and 3 **C** 3 and 4 **D** 4 and 1

34 Scientists have investigated the absorption of mineral ions by plant roots. They believe that it needs energy from respiration.

Which observation best supports this idea?

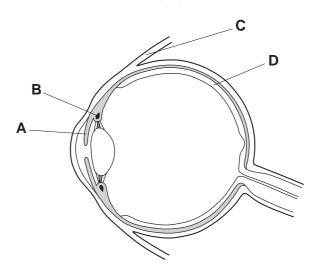
- A Carbohydrate is stored in the roots.
- **B** Living roots give off carbon dioxide.
- **C** Nitrate uptake is reduced in lower oxygen concentrations.
- **D** The root hairs have a large surface area for gas exchange.
- **35** The diagram represents the blood supply to the liver and kidneys.

Which vessel contains blood with the lowest concentration of urea?



36 The diagram shows an eye in section.

Which structure is mainly responsible for changing the size of the pupil?



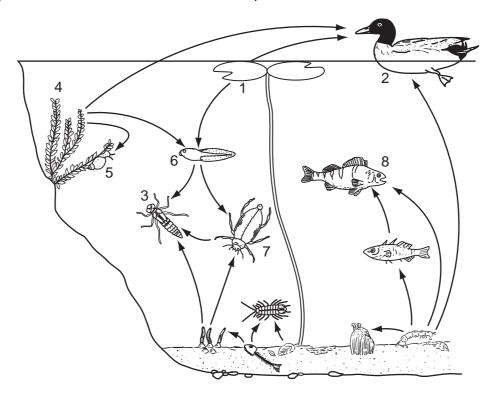
37 The diagram shows the label from a bottle of gin.



What will happen, during the next few hours, after a person drinks a large amount of gin?

- **A** Their judgement of distance will improve.
- **B** Their muscle control will be reduced.
- **C** Their reaction time will decrease.
- **D** Their urine output will decrease.

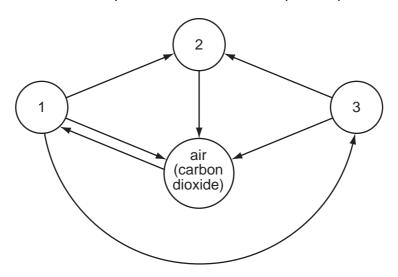




Which of the organisms is a producer, a herbivore or a carnivore?

	producer	herbivore	carnivore	
Α	1	6	7	
В	2	4	5	
С	4	2	6	
D	7	3	8	

39 In the diagram, arrows represent the movements of carbon compounds in the carbon cycle. The circles represent carbon compounds in animals, decomposers, plants and in the air.



What is represented by each circle?

	1	2	3		
Α	decomposers	animals	plants		
В	animals	decomposers	plants		
С	plants	decomposers	animals		
D	decomposers	plants	animals		

40 Where does the exchange of materials take place between mother and fetus?

- **A** oviduct
- **B** umbilical cord
- C uterus
- **D** vagina

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DATA SHEET
The Periodic Table of the Elements

	0	Heium Heium	20 Neon 10 Ar Argon	84 Krypton 36	131 Xe Xenon 54	Radon 86		Lutetium	Lr Lawrencium 103														
	II/		19 Fluorine 9 35.5 C1 Chlorine	80 Brownine 35	127 I lodine 53	At Astatine 85		173 Yb Ytterbium 70	No Nobelium 102														
	I					16 Oxygen 8 32 Sulphur 16	Selenium 34	128 Te Tellurium	Po Polonium 84		169 Tm Thulium 69	Md Mendelevium 101											
	>		Nitrogen 7 31 Phosphorus 15	75 AS Arsenic	Sb Antimony 51	209 Bi Bismuth 83		167 Er Erbium 68	Fm Fermium 100														
	<u>></u>		12 Carbon 6 Carbon 8 Si Siicon 14	73 Ge Germanium 32	119 Sn Tin 50	207 Pb Lead		165 Ho Holmium 67	Esteinium 99														
	≡		11 B Boron 5 27 A1 Auminium 13	70 Ga Gallium 31	115 In Indium 49	204 T 1 Thallium 81		162 Dy Dysprosium 66	Cf Californium 98														
			·	65 Zn Zinc 30	Cd Cadmium 48	201 Hg Mercury 80		159 Tb Terbium 65	BK Berkelium 97														
				64 Copper 29	108 Ag Silver 47	197 Au Gold		Gd Gadolinium 64	Curium 96														
Group		Hydrogen																59 X Nickel 28	106 Pd Palladium 46	195 Pt Platinum 78		152 Eu Europium 63	Am Americium 95
Ğ				59 Co Cobalt	103 Rh Rhodium 45	192 Ir Iridium 77		Sm Samarium 62	Pu Plutonium 94														
			Hydrogen		56 Fe Iron	Ruthenium 44	190 Os Osmium 76		Pm Promethium 61	Np Neptunium 93													
					Mn Manganese 25	Tc Technetium 43	186 Re Rhenium 75		Neodymium 60	238 U Uranium 92													
				52 Cr Chromium 24	96 Mo Molybdenum 42	184 W Tungsten 74		Pr Praseodymium 59	Pa Protactinium 91														
				51 V Vanadium 23	Niobium 41	181 Ta Tantalum 73		140 Ce Cerium	232 Th Thorium														
				48 Ti Titanium	91 Zr Zirconium 40	178 Hf Hafnium 72			nic mass Ibol nic) number														
				Scandium 21	89 Y Yttrium 39	139 La Lanthanum 57 *	227 Ac Actinium 89	d series series	a = relative atomic mass X = atomic symbol b = proton (atomic) number														
	=		Be Berylium 4 24 Magnesium 12	40 Ca Calcium	Strontium	137 Ba Barium 56	226 Ra Radium 88	*58-71 Lanthanoid series	т х в														
	_		7 Lithium 3 23 Na Sodium 11	39 Potassium	85 Rb Rubidium 37	Caesium 55	Fr Francium 87	*58-71 L	Key														

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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