

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

General Certificate of Education O Level

MARK SCHEME for the November 2004 question paper

5070 CHEMISTRY

5070/04

Paper 4 (Alternative to Practical), maximum mark 60

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NOVEMBER 2004

GCE O Level

MARK SCHEME

MAXIMUM MARK: 60

SYLLABUS/COMPONENT: 5070/04

CHEMISTRY
Paper 4 (Alternative to Practical)



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- 1 24(1) cm³ [1]
- 2 (a) (i) ethanol (1), C₂H₅OH (1) (e.c.f. allowed or mark separately for ethanol or correct formula)
- (b) yeast (1)
- (c) when the thermometer showed an increase, or temperature rises above the boiling point of ethanol (1)
- not no more distillate produced.
- (d) (i) orange (1) to green (1) (mark separately within reason)
- (ii) ethanoic acid (1)
- (e) (i) ethyl ethanoate (1), CH₃COOC₂H₅ (1) no e.c.f except for an ester. Not C₄H₈O₂
- (ii) esters (1) e.c.f allowed here from (e) (i)
- (carboxylic acids not organic acids if appropriate e.c.f.) [10]
- 3 (a) chromatography (1)
- (b) line drawn below base line (1) (must be straight, using a ruler, and parallel with the base line)
- (c) ink consists of different colours, dyes, components (1) which would be separated (1)
- (d) X contains S and U. (1)
- Y contains R, S, and T. (1) (all in each case for 1 mark)
- (e) distance travelled by t = 4 cm
- distance travelled by solvent front = 5.5 cm (both 1)
- (No other values are acceptable as they are drawn exactly at 4.0 and 5.5 cm.)
- R_f = 4/5.5 = 0.72 or 0.73 (1) (to two d.p.) (not 0.7)
- Accept any e.c.f. for R_f even if >1 [8]
- Question 4 to 7 (a), (b), (c). 1 mark each** [4]
- 8 (a) 1.55 g (1)
- (b) to allow gas to escape (1) or to prevent pressure build up.
- (c) red or pink or orange to yellow (1)

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- (d) 24.1 41.1 28.5 1 mark for
0.0 17.6 4.8 correct row or
24.1 23.5 23.7 column (3)

mean value = 23.6 (1) cm³

- (e) 0.00236 (1) (i) 0.0264 (1)
(f) 0.00236 (1) (j) 0.0132 (1)
(g) 0.0236 (1) (k) (i) 100 g (1)
(h) 0.05 (1) (ii) 1.32 g (1)
(iii) 85.2% (1)

[16]

9 1 colourless solution (1) (no substances or solids.)

2 (a) white ppt. (1)

(b) soluble in excess (1)

3 (a) white ppt. (1)

(b) insoluble in excess (1)

4 Al foil (1), aq NaOH and heat (1), NH₃ or
gas evolved (1), test for NH₃ (1).

Al foil (1) followed by incorrect chemistry loses the second mark and the ammonia or gas evolved mark. The test for ammonia may be scored if correct.

or 'Brown Ring' test:

aq. FeSO₄ (1), conc. H₂SO₄ (1), aq. and conc. (1)

brown ring (1)

Formula Al(NO₃)₃ (1)

[10]

10 (a) hydrogen (1)

(b) 18, 40, 54, 60 (2) all correct (one error 1 mark)

(c) all points, recorded in the table, plotted correctly (1) two smooth curves (1), any attempt to draw reasonable curves (no straight lines between points) both passing through zero (1)

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(d) (i) 48 (1) cm³

(ii) 2.6 (1) minutes (in both cases read candidates graph and insist **to** half a small square)

(e) (i) powdered (1)

25 cm³ of 0.200 mol/dm³ or equivalent (2 or 0)

or double the concentration and halve the volume (2)

or increase the concentration and reduce the volume to give the same number of moles (2)

(increase concentration and reduce volume worth 1)

[11]