



Cambridge IGCSE™ (9–1)

CANDIDATE
NAME

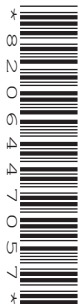
--

CENTRE
NUMBER

--	--	--	--	--

CANDIDATE
NUMBER

--	--	--	--



CO-ORDINATED SCIENCES

0973/31

Paper 3 Theory (Core)

May/June 2020

2 hours

You must answer on the question paper.

No additional materials are needed.

INSTRUCTIONS

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

INFORMATION

- The total mark for this paper is 120.
- The number of marks for each question or part question is shown in brackets [].
- The Periodic Table is printed in the question paper.

This document has **32** pages. Blank pages are indicated.

- 1 (a) Fig. 1.1 is a diagram of the human female reproductive system.

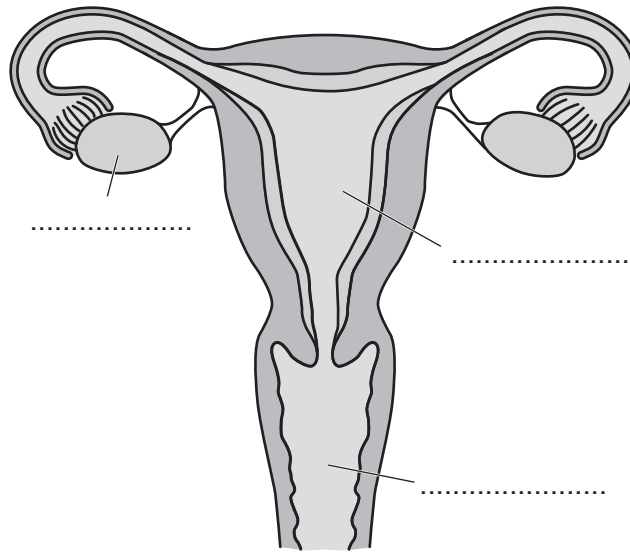


Fig. 1.1

- (i) Label Fig. 1.1 using words from the list.

Each word may be used once, more than once or not at all.

cervix

ovary

oviduct

uterus

vagina

[3]

- (ii) The boxes on the left are parts of the human female reproductive system.

The boxes on the right are the functions of the different parts.

Draw one line from each part to its function.

part

function

ovary

site of fertilisation

oviduct

receives penis during sexual intercourse

uterus

releases female gametes

vagina

where fetus develops

[3]

(b) HIV is a sexually transmitted disease often spread through unprotected sexual intercourse.

Fig. 1.2 is a graph showing the number of new infections of HIV in a country between 1999 and 2009.

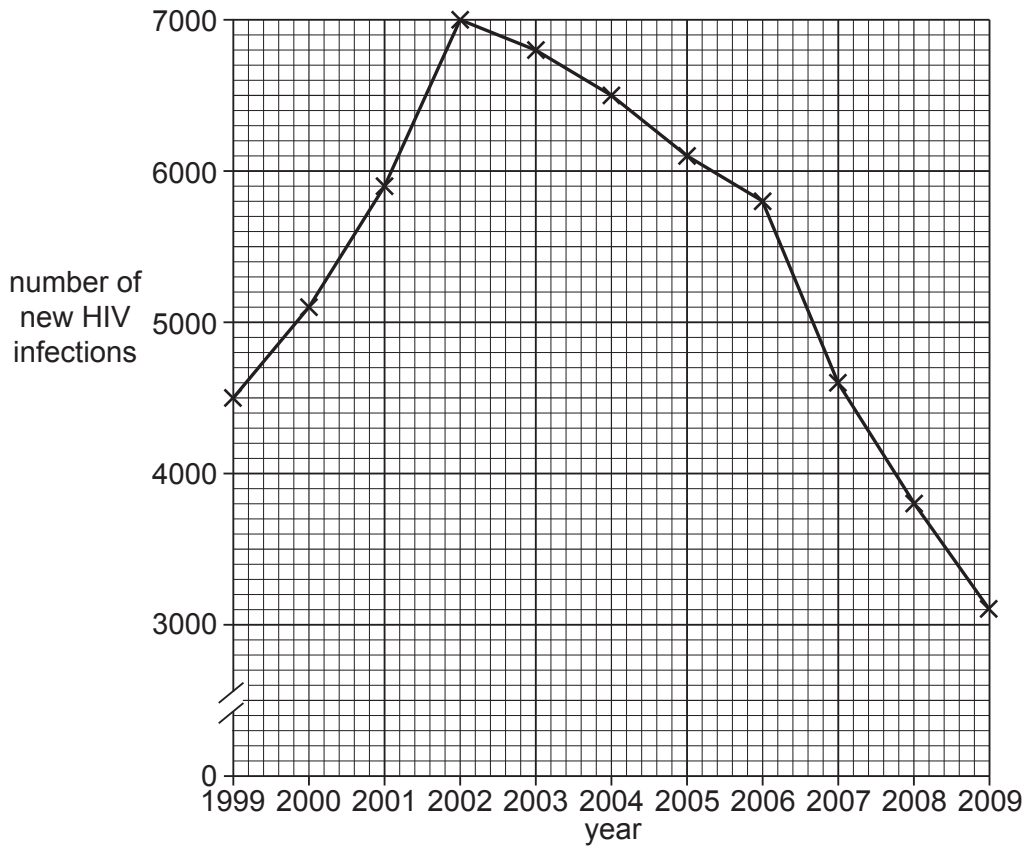


Fig. 1.2

(i) Calculate the overall change in the number of new HIV infections between 1999 and 2009.

..... [1]

(ii) Describe the trends in new HIV infections between 1999 and 2009.

Use data to support your answer.

.....

 [2]

(c) Table 1.1 shows methods of how HIV can be transmitted or controlled.

Complete Table 1.1 to identify if the method is an example of **transmission** or **control**.

Table 1.1

method	transmission or control
barrier contraception	
contaminated blood transfusion	
sharing needles when taking drugs	

[2]

[Total: 11]

- 2 (a) The diagram in Fig. 2.1 shows part of the water cycle.
- Clouds form above the sea and rain falls from the clouds.
- Clouds are made up of very small drops of liquid water.

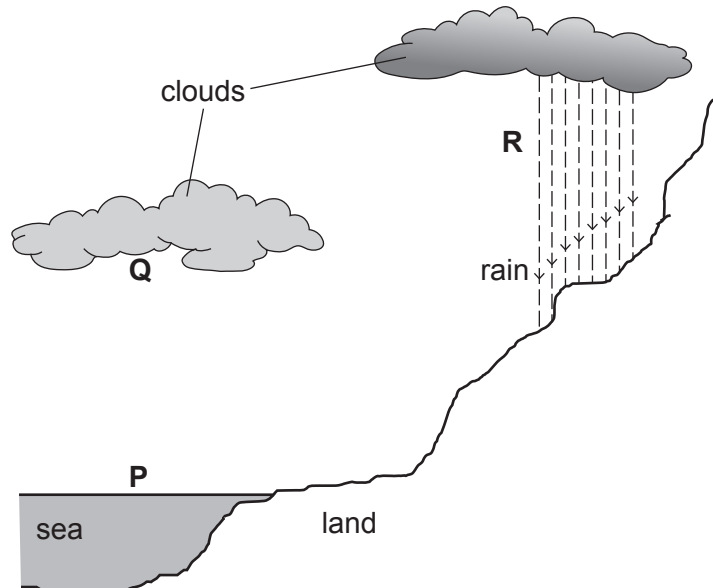


Fig. 2.1

The letters **P**, **Q** and **R** show locations where **physical** changes happen in the water cycle.

- (i) Use the letters **P**, **Q** and **R** to complete Table 2.1.

Each letter may be used once, more than once or not at all.

Table 2.1

description of change	location
water vapour condenses	
liquid water gains kinetic energy	
water molecules move closer together	
water evaporates	

[3]

- (ii) Explain why the changes **P**, **Q** and **R** are **physical** changes.

.....

..... [1]

(b) Carbon dioxide in the air dissolves in rainwater. This causes the rainwater to become slightly acidic.

(i) Water is **neutral**.

State the pH value of water.

.....

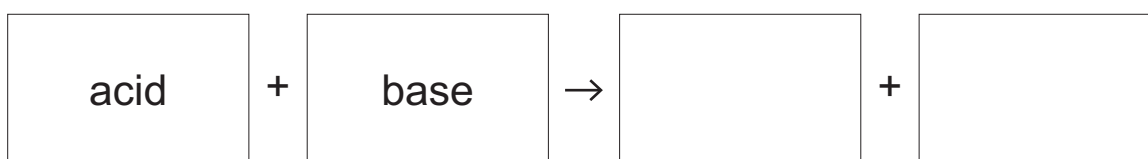
[1]

(ii) Suggest a pH value of rainwater.

.....

[1]

(c) Complete the word equation for the neutralisation reaction between an acid and a base.



[1]

(d) Table 2.2 shows some oxides.

It also shows whether they are acidic or basic.

Table 2.2

oxide	acidic or basic
carbon dioxide	acidic
chlorine oxide	acidic
magnesium oxide	basic
phosphorus oxide	acidic
sodium oxide	basic

(i) Predict whether nitrogen dioxide is acidic or basic.

Explain your answer.

nitrogen dioxide is

explanation

.....

[1]

(ii) Nitrogen dioxide is an air pollutant.

Describe **one** human activity that releases nitrogen dioxide into the air.

.....
..... [1]

(iii) Identify **one** other gaseous air pollutant that is harmful to humans.

..... [1]

[Total: 10]

- 3 (a) Fig. 3.1 shows a skier at the top of a slope.



Fig. 3.1

- (i) The skier travels 310.5 m in 20 s.

Calculate the average speed of the skier.

average speed = m/s [2]

- (ii) Fig. 3.2 shows a speed-time graph of the skier.

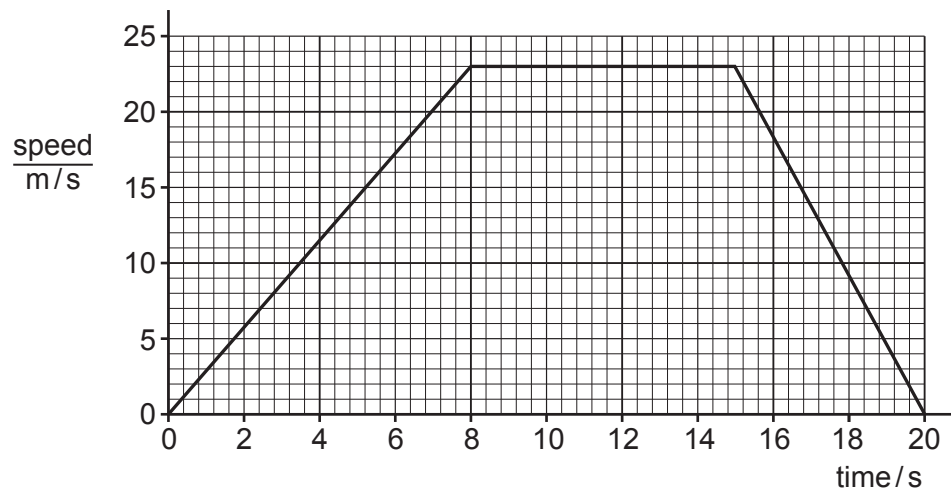


Fig. 3.2

Use Fig. 3.2 to determine the distance travelled while the skier has constant acceleration during the first 8 seconds.

distance = m [2]

(iii) State the name of the force that impedes the skier's motion.
..... [1]

(b) (i) The skier has a mass of 85 kg.
The gravitational field strength g is 10 N/kg.
Calculate the weight of the skier.

weight = N [2]

(ii) State the source of the gravitational field that causes the skier to accelerate down a slope.
..... [1]

(iii) When the skier goes to the top of a slope, he does work.
As he climbs, his gravitational potential energy increases.
Choose from the list the correct word to complete the sentence below.

- created gained lost transferred**

The work done by the skier is equal to the total energy [1]

[Total: 9]

4 A student investigates photosynthesis using an aquatic plant (Elodea).

Fig. 4.1 shows the apparatus the student uses.

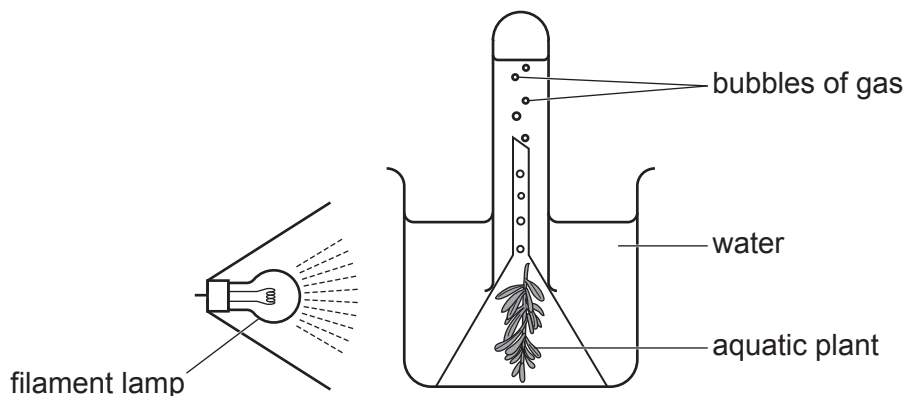


Fig. 4.1

(a) (i) The aquatic plant releases bubbles of gas.

Name the gas the bubbles are made from.

..... [1]

(ii) The investigation was repeated with the filament lamp removed.

Explain why the number of bubbles produced decreases.

.....
 [1]

(b) Photosynthesis is an enzyme-controlled reaction.

(i) Define the term *enzyme*.

.....
 [2]

(ii) Suggest why placing a hot filament lamp too close to the water could stop the plant producing bubbles.

.....
 [1]

(c) Fig. 4.2 is a photomicrograph of the cells in an Elodea leaf.

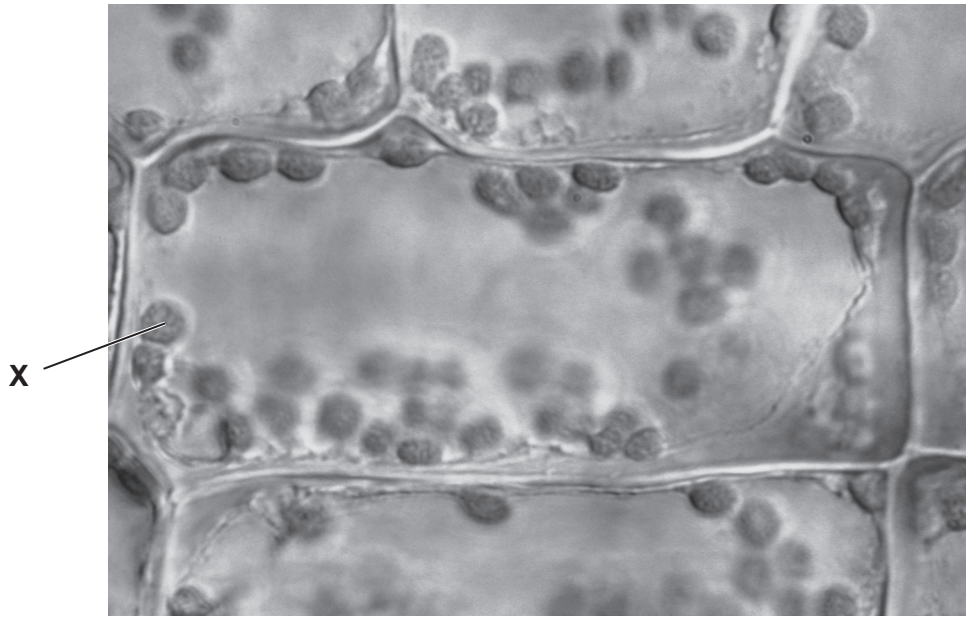


Fig. 4.2

(i) Identify the part of the cell labelled **X** in Fig. 4.2.

..... [1]

(ii) Name three parts of a cell found in **both** plant cells and animal cells.

1

2

3

[3]

(iii) Suggest **two** ways the structure of root hair cells would differ from the cells shown in Fig. 4.2.

.....

..... [2]

(iv) State **one** function of root hair cells.

.....

..... [1]

[Total: 12]

5 Non-metallic elements are placed in groups towards the right hand side of the Periodic Table.

(a) Describe what is meant by a group in the Periodic Table.

.....
 [1]

(b) Fig. 5.1 shows the chemical symbols of five non-metallic elements.

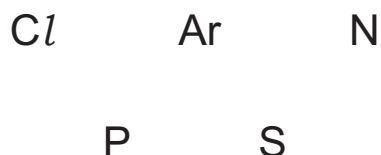


Fig. 5.1

(i) One of the symbols in Fig. 5.1 is **not** in the same **period** of the Periodic Table as the others.

Identify the symbol and explain your answer.

symbol

explanation

.....

..... [1]

(ii) Select **two** elements from Fig. 5.1 that are contained in fertilisers.

..... and [1]

(iii) State **one** reason why fertiliser is added to soil.

.....

..... [1]

(c) Select **one** element from Fig. 5.1 that is used to treat water to make it safe to drink.

Explain how it does this.

element

explanation

..... [2]

(d) Fig. 5.2 shows the electronic structures of atoms of argon, chlorine and potassium.

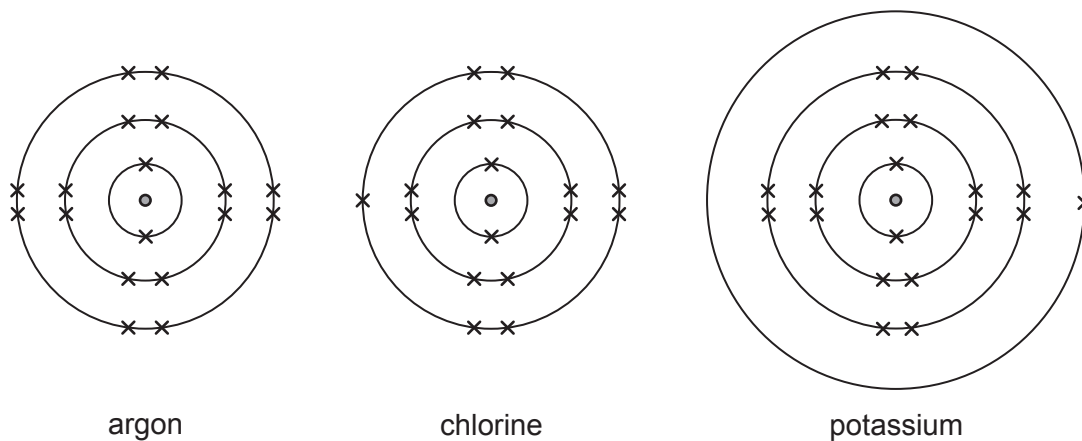


Fig. 5.2

(i) Argon does **not** react with potassium.

Explain why.

Use ideas about electronic structures in your answer.

.....

.....

..... [2]

(ii) Potassium reacts with chlorine to form potassium chloride.

Describe how the electronic structures of a potassium atom and of a chlorine atom change when potassium and chlorine react.

change in potassium

.....

change in chlorine

.....

[2]

(iii) Name the type of chemical bonding in potassium chloride.

..... [1]

[Total: 11]

6 (a) Solar cells can be used to generate electricity for a house.

State one advantage and one disadvantage, apart from cost, of solar cells.

advantage

disadvantage

[2]

(b) Fig. 6.1 shows an ice cube and a thermometer in a glass of water.

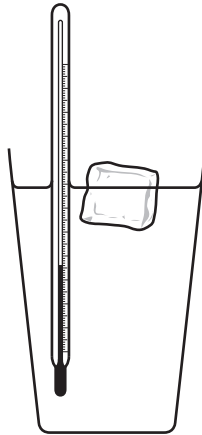


Fig. 6.1

(i) The water provides thermal energy which melts the ice.

State what happens to the temperature of the ice as it is melting.

..... [1]

- (ii) Fig. 6.2 shows that more ice cubes have been added to the glass of water.

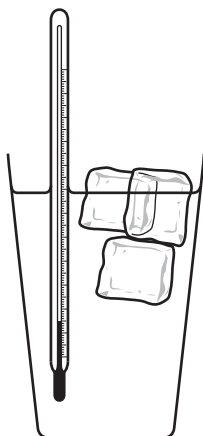


Fig. 6.2

The ice cubes reduce the temperature of the water. The scale on the liquid-in-glass thermometer shows this decrease in temperature.

Explain this action of the liquid-in-glass thermometer.

Use ideas about particles in your answer.

.....

.....

..... [2]

- (iii) The water in the glass is evaporating.

Describe evaporation in terms of the motion of the water molecules.

.....

..... [2]

[Total: 7]

7 (a) A teacher measured the height of the students in a class to the nearest cm.

Table 7.1 shows the results.

Table 7.1

height range/cm	frequency in class
140–145	2
146–150	6
151–155	8
156–160	12
161–165	5
166–170	2
171–175	1

(i) State the most common height range in the class.

..... cm [1]

(ii) State the type of variation shown by height.

..... [1]

(iii) State the evidence from Table 7.1 that supports your answer in (a)(ii).

.....
 [1]

(b) Use words from the list to complete the sentences about natural selection.

Each word may be used once, more than once or not at all.

- alleles
 - cells
 - competition
 - die
-
- react
 - survive
 - variety

Individuals in a population produce offspring.

If there are not enough resources for all the individuals there is increased

.....

Individuals that are better adapted to the environment

Individuals that are not adapted

Individuals that can reproduce passing their alleles to the next generation.

[4]

(c) Lack of resources can cause extinction of a species.

Suggest two ways humans can cause extinction.

1

2

[2]

[Total: 9]

8 Calcium and copper are metals.

(a) (i) State two **physical** properties of metals.

- 1
- 2 [2]

(ii) Copper is a transition metal. Calcium is **not** a transition metal.

State two properties of transition metals that are **not** properties of calcium.

- 1
-
- 2
- [2]

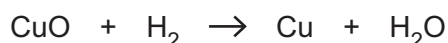
(iii) Calcium and copper are added to separate amounts of water.

Describe the reaction, if any, for each metal.

- calcium
- copper [2]

(b) Copper metal is produced when copper(II) oxide reacts with hydrogen gas.

The equation for this reaction is shown.



(i) Explain why this equation is described as *balanced*.

-
- [1]

(ii) Identify which substance is **reduced** in this reaction.

Explain your answer.

- substance
- explanation
- [2]

(c) Metals are covered with a thin layer of copper during electroplating.

Fig. 8.1 shows apparatus and materials to electroplate a metal key with copper.

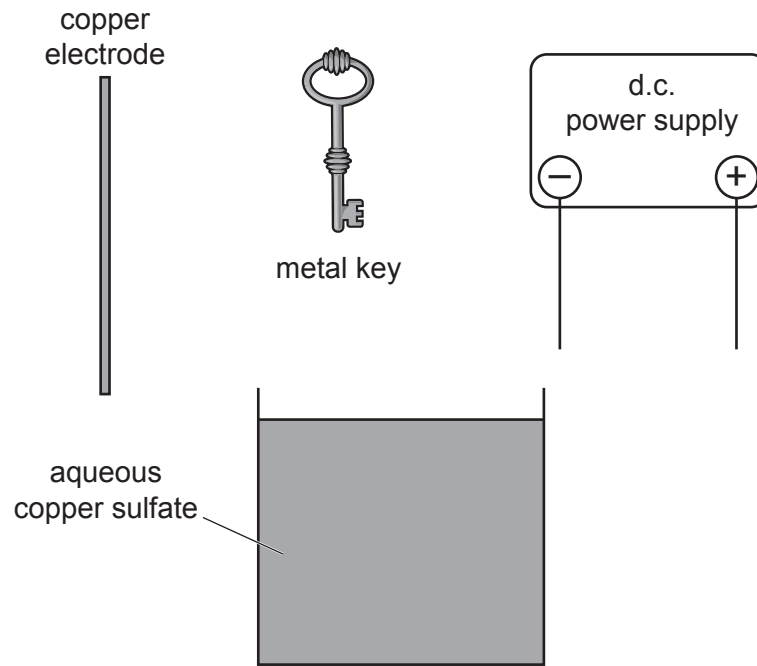


Fig. 8.1

Draw a diagram to show the apparatus and materials set up to electroplate the metal key with copper.

[2]

[Total: 11]

- 9 (a) Fig. 9.1 shows a thin converging lens used in a digital camera.

A ray of light has been drawn from a man's head to the image sensor.

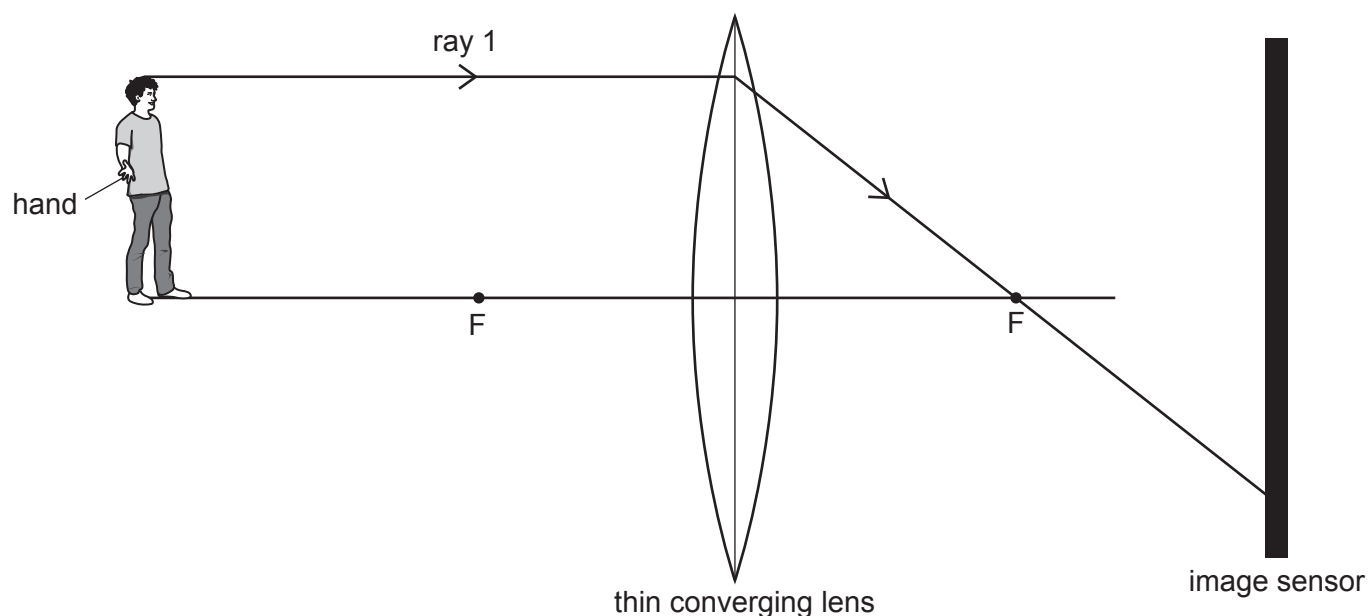


Fig. 9.1

- (i) On Fig. 9.1 draw a ray of light from the man's **hand** to show where it will be detected on the image sensor of the camera. [2]
- (ii) The image is formed on the image sensor.

Circle the **two** correct words or phrases that describe the image.

diminished

enlarged

inverted

same size

upright

[2]

- (iii) The camera detects visible light, and has an infrared sensor.

Write **visible light** and **infrared** in the correct positions in the electromagnetic spectrum in Fig. 9.2.



[2]

Fig. 9.2

- (b) The camera is used to photograph a thunder storm.

Thunder and lightning are caused at the same time. The photographer sees the flash of lightning before he hears the thunder.

- (i) Explain why the photographer sees the lightning before he hears the thunder.

.....
 [1]

- (ii) Explain why an astronaut orbiting the Earth in a space-station sees the lightning but does **not** hear the thunder.

.....

 [2]

(c) When electronic equipment is recycled, some of the materials can be sorted using magnets.

(i) In a recycling factory an electromagnet is used to sort steel from other metals.

Explain why an electromagnet is used to sort the steel.

.....

.....

..... [2]

(ii) Some materials at the recycling factory were tested to see if they conducted electricity.

Complete Table 9.1 by placing a tick (✓) in the electrical conductor column or electrical insulator column to correctly describe each material.

Table 9.1

	electrical conductor	electrical insulator
aluminium		
cardboard		
copper		
polystyrene		
PVC		

[2]

[Total: 13]

10 Fig. 10.1 is a drawing of a person doing a parachute jump.

Before the jump the person experiences a large increase in the production of the hormone adrenaline, which targets the liver and the heart.



Fig. 10.1

(a) Describe how adrenaline is transferred to the heart and the liver.

.....

..... [1]

(b) Place ticks (✓) in the boxes to show **all** the effects of adrenaline on the body in Table 10.1.

Table 10.1

decreased pulse rate	
increased breathing rate	
increased transpiration	
mutation of DNA	
widened pupils	

[2]

(c) One other effect of adrenaline is an increased rate of respiration.

(i) State the word equation for respiration.

..... [2]

(ii) State why respiration is needed for muscle contraction.

.....
..... [1]

(d) Respiration is one of the characteristics of living things.

State two **other** characteristics of living things.

1

2 [2]

[Total: 8]

11 Petroleum and natural gas are extracted from the Earth.

(a) Gasoline is a useful product made from petroleum by fractional distillation.

(i) State a use for gasoline.

..... [1]

(ii) State **one other** useful product made from petroleum by fractional distillation.

..... [1]

(b) Compound **G** is the main constituent of natural gas.

(i) State the name of compound **G**.

..... [1]

(ii) Fig. 11.1 is a diagram of one molecule of compound **G**.

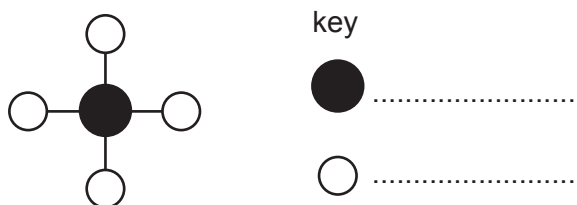


Fig. 11.1

On Fig. 11.1, complete the key to identify the atoms in the molecule of **G**. [1]

(c) Alkanes are **saturated** compounds.

Alkenes are **unsaturated** compounds.

(i) Describe a test and its results that shows whether a compound is an alkane or an alkene.

test

result with an alkane

.....

result with an alkene

.....

[2]

(ii) Describe the difference in the covalent bonding in alkanes and in alkenes.

.....

..... [1]

(iii) State the name of the chemical reaction that makes alkenes from alkanes.

..... [1]

[Total: 8]

12 (a) Fig. 12.1 shows a circuit containing a battery of 4 cells.

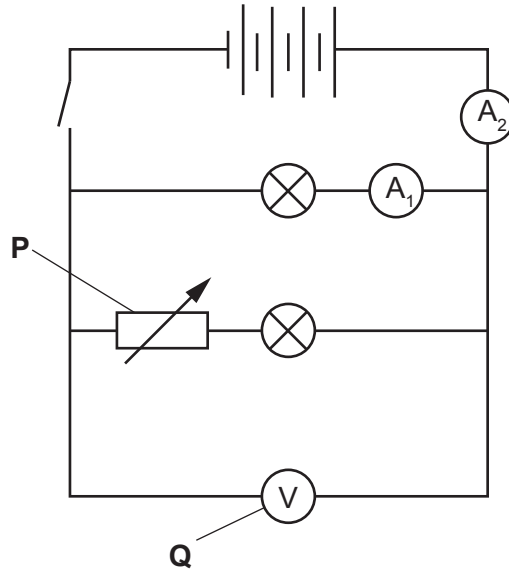


Fig. 12.1

(i) Name the components P and Q.

component P

component Q

[2]

(ii) The battery is a source of electromotive force (e.m.f.).

State the unit of e.m.f.

unit =

[1]

(iii) The switch is closed and both lamps light up.

Readings are recorded on ammeters A₁ and A₂.

Describe the difference, if any, in the readings of A₁ and A₂.

Explain your answer.

difference

explanation

.....

[2]

(b) Fig. 12.2 shows a mains operated d.c. power source.

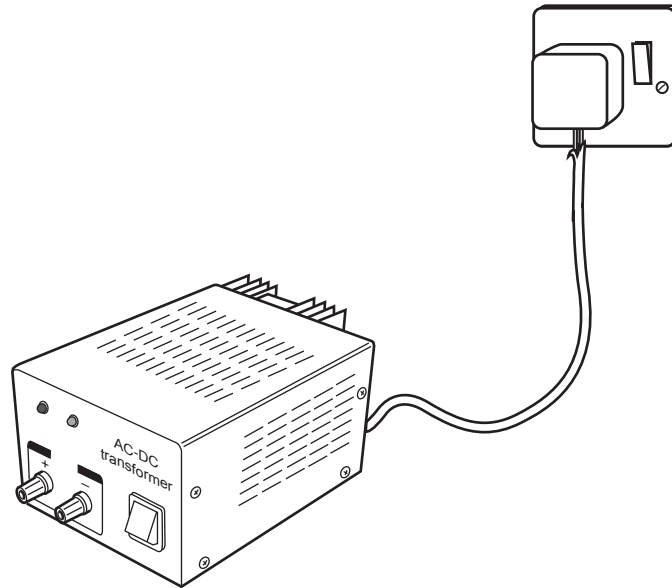


Fig. 12.2

Identify one electrical hazard on Fig. 12.2.

..... [1]

(c) Argon gas is used in some types of lamp.

An argon atom has the chemical symbol ${}^{40}_{18}\text{Ar}$.

State the composition of the nucleus of an atom of Argon.

.....

 [2]

(d) A sample of radioactive rock was tested to see if it emitted α -particles.

(i) Describe how a radiation detector could be used to show that α -particles were being emitted.

.....
.....
.....
..... [2]

(ii) When the sample of radioactive rock is removed from the detector, the detector continues to record some radiation.

Explain this observation.

.....
..... [1]

[Total: 11]

BLANK PAGE

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

To avoid the issue of disclosure of answer-related information to candidates, all copyright acknowledgements are reproduced online in the Cambridge Assessment International Education Copyright Acknowledgements Booklet. This is produced for each series of examinations and is freely available to download at www.cambridgeinternational.org after the live examination series.

Cambridge Assessment International Education is part of the Cambridge Assessment Group. Cambridge Assessment is the brand name of the University of Cambridge Local Examinations Syndicate (UCLES), which itself is a department of the University of Cambridge.

The Periodic Table of Elements

		Group							
I	II	III	IV	V	VI	VII	VIII		
3 Li lithium 7	4 Be beryllium 9	1 H hydrogen 1	5 B boron 11	6 C carbon 12	7 N nitrogen 14	8 O oxygen 16	9 F fluorine 19	10 Ne neon 20	
11 Na sodium 23	12 Mg magnesium 24	Key atomic number atomic symbol name relative atomic mass		13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	18 Ar argon 40
19 K potassium 39	20 Ca calcium 40	21 Sc scandium 45	22 Ti titanium 48	23 V vanadium 51	24 Cr chromium 52	25 Mn manganese 55	26 Fe iron 56	27 Co cobalt 59	28 Ni nickel 59
37 Rb rubidium 85	38 Sr strontium 88	39 Y yttrium 89	40 Zr zirconium 91	41 Nb niobium 93	42 Mo molybdenum 96	43 Tc technetium —	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106
55 Cs caesium 133	56 Ba barium 137	57–71 lanthanoids	72 Hf hafnium 178	73 Ta tantalum 181	74 W tungsten 184	75 Re rhenium 186	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195
87 Fr francium —	88 Ra radium —	89–103 actinoids	104 Rf rutherfordium —	105 Db dubnium —	106 Sg seaborgium —	107 Bh bohrium —	108 Hs hassium —	109 Mt meitnerium —	110 Ds darmstadtium —
							80 Hg mercury 201	79 Au gold 197	81 Tl thallium 204
							82 Pb lead 207	83 Bi bismuth 209	84 Po polonium —
							112 Cn copernicium —	111 Rg roentgenium —	114 Fl flerovium —
									116 Lv livermorium —

lanthanoids

57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium —	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
89 Ac actinium —	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	97 Bk berkelium —	98 Cf californium —	99 Es einsteinium —	100 Fm fermium —	101 Md mendelevium —	102 No nobelium —	103 Lr lawrencium —

actinoids

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).