

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

International General Certificate of Secondary Education

MARK SCHEME for the November 2004 question paper

0620 CHEMISTRY

0620/03

Paper 3 (Extended Theory), maximum mark 150

This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which Examiners were initially instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began. Any substantial changes to the mark scheme that arose from these discussions will be recorded in the published *Report on the Examination*.

All Examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes must be read in conjunction with the question papers and the *Report on the Examination*.

- CIE will not enter into discussion or correspondence in connection with these mark schemes.

CIE is publishing the mark schemes for the November 2004 question papers for most IGCSE and GCE Advanced Level syllabuses.



Grade thresholds taken for Syllabus 0620 (Chemistry) in the November 2004 examination.

	maximum mark available	minimum mark required for grade:			
		A	C	E	F
Component 3	150	52	34	25	19

The threshold (minimum mark) for B is set halfway between those for Grades A and C. The threshold (minimum mark) for D is set halfway between those for Grades C and E. The threshold (minimum mark) for G is set as many marks below the F threshold as the E threshold is above it.

Grade A* does not exist at the level of an individual component.



November 2004

INTERNATIONAL GCSE

MARK SCHEME

MAXIMUM MARK: 150

SYLLABUS/COMPONENT: 0620/03

CHEMISTRY
Extended Theory



Page 1	Mark Scheme	Syllabus	Paper
	IGCSE – November 2004	0620	3

- 1 (a) carbon dioxide, water vapour, noble gases **or** a named noble gas
Any **TWO** [2]
- (b) burning fossil fuels [1]
COND that contain sulphur [1]
acid rain **or** any effect of acid rain - deforestation, effect on stone work,
corrosion of metals, acidity in lakes, health etc [1]
- motor vehicles **or** petrol **or** car exhausts [1]
health, if specified then brain, nervous system, development of children etc [1]
do not select from list illnesses
- OR** lead in old paint [1]
harmful effect as above [1]
- (c) (i) combustion **or** burning **NOT** dissolving in the ocean [1]
- (ii) $6\text{CO}_2 + 6\text{H}_2\text{O}$ [1]
exothermic [1]
- (d) (i) glowing splint burst into flame or rekindled [1]
Must have glowing or equivalent idea
OR any similar description that includes the two points glowing and relights.
- (ii) measure volume **or** count bubbles [1]
time [1]
NOT units
- (iii) rate slows down [1]
Because the reaction is photochemical **or** rate depends on intensity of light
or light less bright or less light falling on plant **or** light provides energy for
photosynthesis etc. [1]
- TOTAL = 15**
- 2 dilute
filter
saturated
cool
blue
sulphate [6]
- TOTAL = 6**
- 3 (a) (i) no change in concentration of reagents **or** rates equal [1]
Accept no change in amounts or it is as if the reaction has Stopped
- (ii) back reaction is endothermic **or** the forward reaction is exothermic [1]
Increase in temperature favours the endothermic reaction which is the back
reaction or vice versa. [1]
NB look for correct conclusion re thermicity and comment re position of
equilibrium.

Page 2	Mark Scheme	Syllabus	Paper
	IGCSE – November 2004	0620	3

	(iii)	increased rate	[1]
		because molecules collide more frequently or concentration of molecules is increased or molecules are closer	[1]
		NOT they have more KE	
		increased yield	[1]
		high pressure favours side with few molecules or smaller volume or moves to reduce the pressure	[1]
		this is product side this can be implied	[1]
	(b) (i)	CO ₂ and H ₂ O	[1]
		balanced	[1]
		2CH ₃ OH + 3O ₂ = 2CO ₂ + 4H ₂ O	
	(ii)	methyl ethanoate	[1]
		water	[1]
	(iii)	Methanoic (acid) accept formic acid	[1]
			TOTAL = 13
4	(a) (i)	Correct equation with a more reactive metal	[1]
	(ii)	Electron loss	[1]
	(iii)	Because they can accept electrons or take electrons away from.....	[1]
	(iv)	Silver or silver(I)	[1]
	(b) (i)	increase	[1]
	(ii)	zinc	
		COND and a correct reason - such as it loses electrons more easily or it is more reactive	[1]
		Need both zinc and reason for the mark.	
	(iii)	from the more reactive to the less reactive NOT just from zinc to lead	[1]
			TOTAL = 7
5	(a)	Group II metals will lose 2e	[1]
		Group VI elements will gain 2e	[1]
	(b)	SCl ₂	[1]
		COND 8e around both chlorine atoms	[1]
		8e around sulphur with 2np and 2bp	[1]
		If x and o reversed ignore if this is the only error	
	(c) (i)	Ions cannot move in solid or can move in liquid	[1]
	(ii)	No ions in sulphur chloride or it is covalent or only molecules or only strontium chloride has ions	[1]
			TOTAL = 7

Page 3	Mark Scheme	Syllabus	Paper
	IGCSE – November 2004	0620	3

- 6 (a) (i) correct structure
 $\text{CH}_2=\text{CCl}_2$ [1]
- (ii) because it has a lower M_r or density or its molecules move faster [2]
it is lighter ONLY [1]
only comment - smaller molecules [0]
answer implies or states sieve idea then [0]
- (b) (i) ester linkage [1]
COND polymer chain showing different monomers and continuation [1]
 $-\text{OOC}-\text{C}_6\text{H}_4-\text{COOCH}_2\text{CH}_2\text{O}-$
- (ii) fats or lipids [1]
- (iii) does not decompose easily when heated [1]
accept similar statements
- (c) (i) does not decompose or non-biodegradable shortage of landfill sites or of space visual pollution
poisonous/toxic/harmful gases when burnt
NOT carbon monoxide, sulphur dioxide. If gas named has to be a correct one eg HCl , HCN
dangerous to animals
Any **TWO** [2]
- (ii) conserve petroleum or save energy [1]
NOT cheaper
- TOTAL = 10**
- 7 (a) (i) $\text{Zn}(\text{OH})_2 = \text{ZnO} + \text{H}_2\text{O}$ [2]
reactant [1] products [1]
- (ii) it would melt or it does not decompose or it does not react [1]
NOT no change
- (iii) blue (solid) [1]
to black (solid) [1]
brown gas [1]

Mark consequentially to any error **but not involving simple integers**

There has to be some evidence that the candidate has attempted to work through the calculation and not merely inserted whole numbers.

For example 2, 1, 160 or 1, 0.5, 80

number of moles of $\text{Fe}_2(\text{SO}_4)_3$ = $1/40$ or 0.025

number of moles of Fe_2O_3 formed = $1/40$ or 0.025

mass of iron(III) oxide formed = $0.025 \times 160 = 4\text{g}$

number of moles of SO_3 produced = $3/40$ or 0.075

volume of sulphur trioxide at r.t.p. = 0.075×25

= 1.8dm^3

[5]

TOTAL = 11

Page 4	Mark Scheme	Syllabus	Paper
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- 8 (a) (i) C_6H_{12} [1]
between 60 to 65°C [1]
- (ii) $C_{12}H_{24}$ [1]
COND giving some indication of the method [1]
- (b) add bromine water **or** potassium manganate(VII) [1]
butene it goes from brown/orange/yellow to colourless
or manganate (VII) from pink to colourless [1]
NOT clear
Cyclobutane it remains brown/orange/yellow **or** manganate (VII) stays pink
or no colour change [1]
Accept does not react
Provided colour of reagent somewhere in the answer [3] is possible
- (c) (i) alcohol [1]
- (ii) $CH_3-CH_2-CHCl-CH_3$ [1]
- (iii) $-CH(CH_3)-CH(CH_3)-$ [2]
or any equivalent diagram
[1] for repeat unit and [1] for continuation

TOTAL = 11